



Designation: E2041 – 23

Standard Test Method for Estimating Kinetic Parameters by Differential Scanning Calorimeter Using the Borchardt and Daniels Method¹

This standard is issued under the fixed designation E2041; the number immediately following the designation indicates the year of original adoption or, in the case of revision, the year of last revision. A number in parentheses indicates the year of last reapproval. A superscript epsilon (ϵ) indicates an editorial change since the last revision or reapproval.

1. Scope*

1.1 This test method describes the determination of the kinetic parameters of activation energy, Arrhenius pre-exponential factor, and reaction order using the Borchardt and Daniels² treatment of data obtained by differential scanning calorimetry. This test method is applicable to the temperature range from 170 K to 870 K (–100 °C to 600°C).

1.2 This treatment is applicable only to smooth exothermic reactions with no shoulders, discontinuous changes, or shifts in baseline. It is applicable only to reactions with reaction order $n \leq 2$. It is not applicable to acceleratory reactions and, therefore, is not applicable to the determination of kinetic parameters for most thermoset curing reactions or to crystallization reactions.

1.3 The values stated in SI units are to be regarded as standard. No other units of measurement are included in this standard.

1.4 *This standard does not purport to address all of the safety concerns, if any, associated with its use. It is the responsibility of the user of this standard to establish appropriate safety, health, and environmental practices and determine the applicability of regulatory limitations prior to use.*

1.5 *This international standard was developed in accordance with internationally recognized principles on standardization established in the Decision on Principles for the Development of International Standards, Guides and Recommendations issued by the World Trade Organization Technical Barriers to Trade (TBT) Committee.*

¹ This test method is under the jurisdiction of ASTM Committee E37 on Thermal Measurements and the direct responsibility of Subcommittee E37.01 on Calorimetry and Mass Loss.

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² Borchardt, H.J., Daniels, F., “The Application of Differential Thermal Analysis to the Study of Reaction Kinetics”, *Journal of the American Chemical Society*, Vol 79, 1957, pp. 41–46.

2. Referenced Documents

2.1 ASTM Standards:³

- E473 Terminology Relating to Thermal Analysis and Rheology
- E537 Test Method for Thermal Stability of Chemicals by Differential Scanning Calorimetry
- E698 Test Method for Kinetic Parameters for Thermally Unstable Materials Using Differential Scanning Calorimetry and the Flynn/Wall/Ozawa Method
- E967 Test Method for Temperature Calibration of Differential Scanning Calorimeters and Differential Thermal Analyzers
- E968 Practice for Heat Flow Calibration of Differential Scanning Calorimeters (Withdrawn 2023)⁴
- E1142 Terminology Relating to Thermophysical Properties
- E1445 Terminology Relating to Hazard Potential of Chemicals
- E1641 Test Method for Decomposition Kinetics by Thermogravimetry Using the Ozawa/Flynn/Wall Method
- E1970 Practice for Statistical Treatment of Thermoanalytical Data
- E2781 Practice for Evaluation of Methods for Determination of Kinetic Parameters by Calorimetry and Differential Scanning Calorimetry
- E3142 Test Method for Thermal Lag of Thermal Analysis Apparatus

3. Terminology

3.1 *Definitions*—Specific technical terms used in this test method are defined in Terminologies E473, E1142, and E1445, including *calibration*, *calorimeter*, *differential scanning calorimetry*, *enthalpy*, *peak*, *reaction*, *repeatability*, *reproducibility*, and *slope*.

³ For referenced ASTM standards, visit the ASTM website, www.astm.org, or contact ASTM Customer Service at service@astm.org. For *Annual Book of ASTM Standards* volume information, refer to the standard's Document Summary page on the ASTM website.

⁴ The last approved version of this historical standard is referenced on www.astm.org.

*A Summary of Changes section appears at the end of this standard

4. Summary of Test Method

4.1 A test specimen is heated at a linear rate in a differential scanning calorimeter or other suitable calorimeter through a region of exothermic reaction behavior. The rate of heat evolution, developed by a chemical reaction, is proportional to the rate of reaction. Integration of the heat flow as a function of time yields the total heat of a reaction.

4.2 The Borchardt and Daniels² data treatment is used to derive the kinetic parameters of activation energy, Arrhenius pre-exponential factor, and reaction order from the heat flow and total heat of reaction information obtained in 4.1 (see Section 5).

5. Basis of Methodology

5.1 Kinetic reactions may be modeled with a number of suitable equations. The Borchardt and Daniels² method makes use of the rate equation to describe the dependence of the rate of reaction on the amount of material present.

$$d\alpha/dt = k(T) (1 - \alpha)^n \quad (1)$$

where:

da/dt = reaction rate (min^{-1})
 α = fraction reacted (dimensionless),
 $k(T)$ = rate constant at temperature T (min^{-1}), and
 n = reaction order (dimensionless).

5.2 For a reaction conducted at temperature (T), the rate equation of Eq 1, may be cast in its logarithmic form:

$$\ln[da/dt] = \ln[k(T)] + n \ln[1 - \alpha] \quad (2)$$

This equation has the form of a straight line, $y = mx + b$, where a plot of the logarithm of the reaction rate ($\ln[da/dt]$) versus the logarithm of the fraction remaining $\ln[1 - \alpha]$ yields a straight line, the slope of which is equal to n and the intercept is equal to $\ln[k(T)]$.

5.3 The Borchardt and Daniels² model also makes use of the Arrhenius equation to describe how the reaction rate changes as a function of temperature:

$$k(T) = Z e^{-E/RT} \quad (3)$$

where:

Z = Arrhenius pre-exponential factor (min^{-1}),
 E = Activation energy (J mol^{-1}),
 T = Absolute temperature (K), and
 R = Gas constant ($= 8.314 \text{ J mol}^{-1} \text{ K}^{-1}$).

5.4 The Arrhenius equation Eq 3 also may be cast in its logarithmic form:

$$\ln[k(T)] = \ln[Z] - E/RT \quad (4)$$

The equation has the form of a straight line, $y = mx + b$, (where $y \equiv \ln[k(T)]$, $m \equiv E/R$, $x \equiv 1/T$ and $b \equiv \ln[Z]$) where a plot of the logarithm of the reaction rate constant ($\ln[k(T)]$) versus the reciprocal of absolute temperature ($1/T$) produces a straight line, the slope of which is equal to $-E/R$ and the intercept of which is $\ln[Z]$.

5.5 As an alternate to Eq 2 and 4, the rate and Arrhenius equations may be combined and cast in its logarithmic form:

$$\ln[da/dt] = \ln[Z] + n \ln[1 - \alpha] - E/RT \quad (5)$$

The resultant equation has the form $z = a + bx + cy$ (where $z \equiv \ln[da/dt]$, $\ln[Z] \equiv a$, $b \equiv n$, $x \equiv \ln[1 - \alpha]$, $c \equiv E/R$, and $y \equiv 1/T$) and may be solved using multiple linear regression data treatment.

5.6 The values for da/dt , $(1 - \alpha)$ and T needed to solve Eq 2, Eq 4 and Eq 5, are experimental parameters obtained from a single linear heating rate DSC experiment scanning through the temperature region of the reaction exotherm as shown in Fig. 1.

5.7 Kinetic results obtained by this test method may be compared with those obtained by Test Method E698.

6. Significance and Use

6.1 This test method is useful in research and development.

6.2 The determination of the appropriate model for a chemical reaction or transformation and the values associated with its kinetic parameters may be used in the estimation of reaction performance at temperatures or time conditions not easily tested. This use, however, is not described in this test method.

7. Interferences

7.1 Because of its simplicity and ease of use, the Borchardt and Daniels² method is often the method of choice for characterization of the kinetic parameters of a reaction system. The Borchardt and Daniels method, like all tools used to evaluate kinetic parameters, is not applicable to all cases. The user of this test method is expressly advised to use this test method and its results with caution.

7.2 Tabulated below are some guidelines for the use of the Borchardt and Daniels² method.

7.2.1 The approach is applicable only to exothermic reactions.

NOTE 1—Endothermic reactions are controlled by the kinetics of the heat transfer of the apparatus and not by the kinetics of the reaction.

7.2.2 The reaction under investigation must have a constant mechanism throughout the whole reaction process. In practice, this means that the reaction exotherm upon heating must be smooth, well-shaped (as in Fig. 1) with no shoulders, multiple peaks or discontinuous steps.

7.2.3 The reaction must be n th order. Confirmation of an n th order reaction may be made by an isothermal experiment such as that described in Appendix X1.

7.2.4 Typical reactions which are not n th order and to which Borchardt and Daniels² kinetic may not be applied for predictive purposes include many thermoset curing reactions and crystallization transformations.

7.2.5 The n th order kinetic reactions anticipate that the value of n will be small, non-zero integers, such as 1 or 2. Values of n greater than 2 or that are not simple fractions, such as $1/2 = 0.5$, are highly unlikely and shall be viewed with caution.

7.2.6 The Borchardt and Daniels² method assumes temperature equilibrium throughout the whole test specimen. This means that low heating rates, (that is, $<10 \text{ K/min}$), small specimen sizes ($<5 \text{ mg}$) and highly conductive sealed specimen containers, for example, aluminum, gold, platinum, etc., should be used.