



Designation: D6714 – 16 (Reapproved 2026)

## Standard Test Method for Chromic Oxide in Ashed Wet Blue (Perchloric Acid Oxidation)<sup>1</sup>

This standard is issued under the fixed designation D6714; the number immediately following the designation indicates the year of original adoption or, in the case of revision, the year of last revision. A number in parentheses indicates the year of last reapproval. A superscript epsilon ( $\epsilon$ ) indicates an editorial change since the last revision or reapproval.

### 1. Scope

1.1 This test method covers the determination of chromic oxide in Wet Blue that has been partly or completely tanned with chromium compounds. In general the samples will contain chromium content between 1 % and 5 %, calculated as chromic oxide expressed upon a dry basis otherwise referred to as moisture-free basis (mfb).

1.2 This test method is specific in that it applies to a sample or samples that are the resultant ash following the execution of Test Method [D6716](#).

1.3 This test method does not apply to Wet White.

1.4 The values stated in SI units are to be regarded as standard. No other units of measurement are included in this standard.

1.5 *This standard does not purport to address all of the safety concerns, if any, associated with its use. It is the responsibility of the user of this standard to establish appropriate safety, health, and environmental practices and determine the applicability of regulatory limitations prior to use.* Specific hazards are given in Section 8.

1.6 *This international standard was developed in accordance with internationally recognized principles on standardization established in the Decision on Principles for the Development of International Standards, Guides and Recommendations issued by the World Trade Organization Technical Barriers to Trade (TBT) Committee.*

### 2. Referenced Documents

2.1 *ASTM Standards:*<sup>2</sup>

[D6658 Test Method for Volatile Matter \(Moisture\) of Wet Blue by Oven Drying](#)

[D6659 Practice for Sampling and Preparation of Wet Blue](#)

<sup>1</sup> This test method is under the jurisdiction of ASTM Committee [D31](#) on Leather and is the direct responsibility of Subcommittee [D31.02](#) on Wet Blue.

Current edition approved July 1, 2026. Published July 2026. Originally approved in 2001. Last previous edition approved in 2021 as D6714 – 16 (2021). DOI: 10.1520/D6714-16R26.

<sup>2</sup> For referenced ASTM standards, visit the ASTM website, [www.astm.org](http://www.astm.org), or contact ASTM Customer Service at [www.astm.org/contact](http://www.astm.org/contact). For *Annual Book of ASTM Standards* volume information, refer to the standard's Document Summary page on the ASTM website.

[and Wet White for Physical and Chemical Tests](#)

[D6716 Test Method for Total Ash in Wet Blue or Wet White](#)

[E180 Practice for Determining the Precision of ASTM Methods for Analysis and Testing of Industrial and Specialty Chemicals](#) (Withdrawn 2009)<sup>3</sup>

### 3. Terminology

3.1 *Definitions*—The terms and definitions employed within this test method are commonly used in normal laboratory practice and require no special comment.

### 4. Summary of Test Method

4.1 The perchloric acid method may be applied to the ash obtained in Test Method [D6716](#). The resultant ash is dissolved into acidic solution in a blend of nitric acid and a prepared oxidation mixture consisting of sulfuric and perchloric acids. Once completed, oxidation of all trivalent to hexavalent chromium is executed by controlled heating. Upon dilution, the chromium is indirectly (back) titrated volumetrically with standardized thiosulfate using released iodine as the titrate. The perchloric acid method requires less manipulation than procedures based on fusion of the ash. However care must be taken because of potential hazards in the use of perchloric acid. The perchloric acid method also tends to give low results.

### 5. Significance and Use

5.1 The procedure described is specific for chromium in Wet Blue. Vanadium is the only common interfering element and this is rarely present in quantity. The precision and accuracy of the methods are usually at least as good as the sampling of the Wet Blue itself, and the accuracy of previously performed test methods.

5.2 The chromium content of Wet Blue is related to the degree of tannage obtained, and hence may be a matter for specification in the purchase of Wet Blue. The procedure described provides adequate accuracy for this purpose.

### 6. Apparatus

6.1 *Analytical Balance*, accurate and calibrated to 0.001 g.

6.2 *Erlenmeyer Flasks*, 250 mL capacity or equivalent.

<sup>3</sup> The last approved version of this historical standard is referenced on [www.astm.org](http://www.astm.org).

6.3 *Burette*, 50 mL capacity of suitable calibration grade, minimum calibration of 0.1 mL.

6.4 *Glass Anti-bumping Beads*, or equivalent.

6.5 *Measuring Cylinders*, 50 mL capacity or equivalent.

6.6 *Small Glass Filter Funnel*.

6.7 *Dessicator*, of suitable size and design and charged with fresh dessicant.

6.8 *Weighing Vessels*, of suitable size and design.

6.9 *Drying Oven*, with accurate variable temperature controls.

6.10 *Hot Plate*, with accurate variable temperature controls.

6.11 *Perchloric Acid Hood*.

## 7. Reagents

7.1 *Purity of Reagents*—Analytical Reagent (AR) grade chemicals shall be used in all tests. Unless otherwise indicated, it is intended that all reagents shall conform to specifications of the Committee on Analytical Reagents of the American Chemical Society,<sup>4</sup> where such specifications are available. Other grades may be used, provided it is first ascertained that the reagent is of sufficiently high purity to permit its use without lessening the accuracy of the determination.

7.2 *Purity of Water*—Unless otherwise indicated, reference to water shall be understood to mean distilled water, deionized water, or water of equal purity.

7.3 *Commercial Reagents*—The use of commercially available pre-standardized analytical reagents and solutions is appropriate, providing those reagents and solutions have been prepared in accordance with and conforming to the previously mentioned specifications (see 7.1).

7.4 *Nitric Acid*—( $\text{HNO}_3$ ), 70 % w/w.

7.5 *Perchloric Acid*—( $\text{HClO}_4$ ), 60 % to 62 % w/w. 70 % or 72 % w/w perchloric acid may be substituted, however, storage is somewhat more hazardous.

7.6 *Sulfuric Acid*—( $\text{H}_2\text{SO}_4$ ), 96 % to 98 % w/w.

7.7 *Potassium Iodide*—(KI), 99 % to 100 % purity.

7.8 *Potassium Iodide Solution*—(KI) 10 % w/w. Dissolve 10 g ( $\pm 0.1$  g) of potassium iodide into 100 mL of water.

7.9 *Starch Indicator Solution*—2 % or equivalent, prepared according to accepted procedures available in analytical handbooks.

7.10 *Oxidizing Mixture*—Mix 1666 mL of concentrated sulfuric acid into an appropriate glass container that contains 2500 mL of perchloric acid, using extreme caution while adding. Cool the mixture to room temperature before use.

NOTE 1—The reagents may be added individually according to 10.3.1.

7.11 *Phosphoric Acid*—( $\text{H}_3\text{PO}_4$ ), 40 % v/v. Dilute 45 mL of 85 % phosphoric acid with water to 100 mL.

7.12 *Hydrochloric Acid*—(HCl), 1:1 dilution of 37 % HCl w/w. Used in standardization of thiosulfate solution.

7.13 *Potassium Dichromate*—( $\text{K}_2\text{Cr}_2\text{O}_7$ ). Used in standardization of thiosulfate solution.

7.14 *Sodium Carbonate*—( $\text{Na}_2\text{CO}_3$ ). Used in preparation of thiosulfate solution.

7.15 *Sodium Thiosulfate*—( $\text{Na}_2\text{S}_2\text{O}_3 \cdot 5\text{H}_2\text{O}$ ). Used to prepare thiosulfate solution.

7.16 *Sodium Thiosulfate Standard Solution*—( $\text{Na}_2\text{S}_2\text{O}_3$ ), 0.1N ( $\pm 0.0002$ ). Dissolve 24.85 g of sodium thiosulfate ( $\text{Na}_2\text{S}_2\text{O}_3 \cdot 5\text{H}_2\text{O}$ ) in previously boiled water, add 1.00 g of sodium carbonate ( $\text{Na}_2\text{CO}_3$ ) and dilute to 1000 mL.

7.16.1 *Standardization*—Dry potassium dichromate ( $\text{K}_2\text{Cr}_2\text{O}_7$ ) in an oven at 130 °C for 2 h and cool in a dessicator. Once cool, weigh about 0.2 g ( $\pm 0.0001$  g) of this dry potassium dichromate into a clean 250 mL, glass-stoppered Erlenmeyer flask. Dissolve in 50 mL of water; add 4 mL of hydrochloric acid (HCl, 1:1) and 20 mL of KI solution. Stopper the flask, and allow to stand 5 min in the dark. Titrate with the thiosulfate solution to be standardized. When the color of the solution has faded to a brownish-green, add 2 mL of 2 % starch solution and continue titrating until the deep blue color changes to a clear green. Record the titration. Calculate the normality of the thiosulfate solution as follows:

$$\text{Normality} = A \div (0.04903 \times B)$$

where:

A =  $\text{K}_2\text{Cr}_2\text{O}_7$  used, g, and

B = amount required for titration, mL.

7.16.2 *Shelf Life*—The thiosulfate solution is relatively stable. However, it should be restandardized at least once a month.

7.17 Alternatively, a commercially available pre-standardized sodium thiosulfate analytical solution may be substituted.

## 8. Hazards

8.1 Chemicals used can be harmful, or explosive, or both.

8.2 The improper use of perchloric acid can lead to violent and serious explosions. In general, these can be traced to situations where concentrated perchloric acid has come in contact with organic or easily oxidized materials.

8.3 The exact procedures given must be followed and the digestion, once started, should be kept from possible contact with other organic matter. The digestion should never be allowed to boil dry. The perchloric acid should never be used without accompanying use of nitric and sulfuric acids.

8.4 All spills involving perchloric acid should be flushed with water and a liquid acid neutralizer. Rags, sawdust, and other organic materials should never be used to mop up spilled acid.

<sup>4</sup> ACS Reagent Chemicals, Specifications and Procedures for Reagents and Standard-Grade Reference Materials, American Chemical Society, Washington, DC. For suggestions on the testing of reagents not listed by the American Chemical Society, see *Analar Standards for Laboratory Chemicals*, BDH Ltd., Poole, Dorset, U.K., and the *United States Pharmacopeia and National Formulary*, U.S. Pharmacopeial Convention, Inc. (USPC), Rockville, MD.